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The hybridization of the central carbon

in $\text{CH}_3\text{C}\equiv\text{N}$ and the

bond angle CCN are

a. sp^2 , 180° . b. sp ,

180° . c. sp^2 ,

120° . d. sp^3 ,

109° . 2. Which of

the following

statements about an

sp hybridized

carbon is FALSE?

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a. It is divalent. b. It forms bonds that are linear. c. It has two p orbitals. d.

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(MO) is most evident in its graphical model of the linear

combination of atomic orbitals (LCAO). ... When atoms approach the distances of chemical bonds, molecular ones are formed from atomic orbitals (AO). ...

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