

Chapter 4 Review Arrangement Of Electrons In Atoms

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opposite spin states. The Pauli exclusion principle states that no two electrons in an atom may have the

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Chapters 4 and 5, which deals with electrons and the periodic table. Includes the chapter vocabulary and a few other useful
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CHAPTER 4 REVIEW . Arrangement of Electrons in Atoms . SHORT ANSWER Answer the following questions in the space
provided. 1. d=-- How many quantum numbers are used to describe the properties of electrons in atomic orbitals? (a) 1 (c)
3 (b) 2 (d) 4 . 2. a,-- A spherical electron cloud surrounding an atomic nucleus would best represent

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CHAPTER 4 REVIEW Arrangement of Electrons in Atoms states that a maximum of two electrons can occupy a single atomic orbital but only if the electrons have opposite spins Hund's rule states that single electrons with the same spin must occupy each equal-energy orbital before additional electrons with opposite spins can occupy the same orbitals

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